

the positions of the atoms around the points  $y=1, z=0$  and  $y=0, z=\frac{1}{2}$  in Fig. 3 shows greater overcrowding in the former case (Table 5). The distance between the methyl at C(13) of the primary molecule ( $x, y, z$ ) and O(4) of the molecule at  $(\bar{x}, \frac{1}{2}+y, \bar{z})$  is 3.94 Å and the distance between the same methyl and O(1) of the molecule at  $(1-x, \frac{1}{2}+y, \bar{z})$  is 3.79 Å. This methyl is held closely between these oxygen atoms. The ring atom

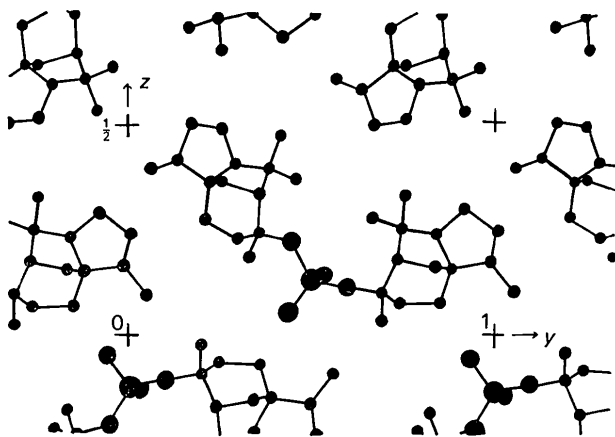


Fig. 3. Molecular packing viewed along the  $a$  axis.

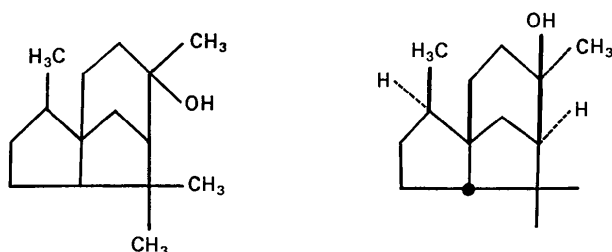


Fig. 4. (+)-cedrol.

C(6) moves into a position to preserve the tetrahedral arrangement at each carbon atom and becomes *trans* to the methyl C(14).

### The absolute configuration

Provided precautions are taken in the indexing of the Weissenberg photographs (Peerdeman & Bijvoet, 1956; Raman, 1958; Ramaseshan, 1964) the structure as determined by the anomalous dispersion method will be in the absolute configuration. Fig. 1 then shows the absolute configuration of the molecule of cedryl chromate and that of (+)-cedrol, deduced from it, is shown in Fig. 4. This is the same as that deduced by Stork & Clarke (1961) from chemical studies.

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## The Crystal and Molecular Structure of Difluoromalonic Acid

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Difluoromalonic acid crystallizes in the monoclinic space group  $P2_1/c$  with cell dimensions  $a=9.59$ ,  $b=5.61$ ,  $c=9.80$  Å,  $\beta=115.1^\circ$  and  $Z=4$ . The crystal is a racemate of two conformational antipodes. The acid molecules in the infinite chains are connected by hydrogen bonds between carboxyl groups across centres of symmetry.

### Introduction

In the crystal structure of  $\alpha$ -substituted monocarboxylic acids, such as propionic acid (Strieter & Templeton, 1962), fluoroacetic acid (Kanters, 1972), glycollic

acid (Pijper, 1971) and several amino-acids (*e.g.* Kanters, Kroon, Beurskens & Vliegthart, 1966), it is found that the carboxylic acid group is nearly coplanar with the  $\alpha$ -substituent. It appears that such a planar arrangement is favoured. These findings are in accord-

ance with theoretical calculations (van Eijck, 1966). The influence of a second carboxyl group in the  $\alpha$ -position upon the conformation can be seen in Table 1.

Two essential cases can be discerned:

(i) the  $\alpha$ -substituent is nearly coplanar with one carboxyl group, the other carboxyl group being almost perpendicular to the first;

(ii) both carboxyl groups strive for coplanarity with the plane through C(1)C(2)R or C(3)C(2)R, complete simultaneous coplanarity being impeded by intramolecular contacts O(2)–O(4) (Table 1).

As far as we know the only disubstituted malonic acid derivative whose structure has been determined is dimethylmalonic acid (Haas & Brenner, 1966). The molecule has a twofold rotation axis. Here no coplanarity of the substituent with the carboxyl group is present, this group is rotated  $30^\circ$  out of the C(1)C(2)R plane.

The coplanarity is most pronounced in monofluoromalonic acid (Roelofsen, Kanters, Kroon & Vliegert, 1971). Therefore one would predict a marked preference for a coplanar arrangement of both halves of the molecule of difluoromalonic acid. In order to check the validity of this assumption we undertook the structure analysis of difluoromalonic acid.

### Experimental

Difluoromalonic acid was synthesized from the dimethyl ester. Very highly hygroscopic crystals (with dimensions  $0.2 \times 0.3 \times 0.8$  mm) were obtained by sublimation *in vacuo*. They were sealed in a Lindemann glass capillary tube. Cell dimensions measured on a General Electric diffractometer were  $a = 9.59 \pm 0.01$ ,  $b = 5.61 \pm 0.01$ ,  $c = 9.80 \pm 0.01$  Å and  $\beta = 115.1 \pm 0.1^\circ$  [ $\lambda(\text{Mo } K\alpha_1\alpha_2) = 0.71069$  Å]. From the systematic extinctions the space group  $P2_1/c$  was deduced. One independent molecule in the asymmetric unit corresponds to the calculated density of  $1.96 \text{ g.cm}^{-3}$ . The intensities were measured on an automatic Nonius three-circle diffractometer equipped with a monochromator, a scintillation counter and a discriminator. The radiation employed was Mo  $K\alpha$ , and the  $\omega$ -scan technique was applied. The 1493 non-zero independent structure amplitudes were

corrected for the Lorentz-polarization factors; no absorption correction was applied.

### Determination of the structure

With Wilson's (1942) method an overall temperature factor ( $B = 2.8 \text{ \AA}^2$ ) and the scaling factor were determined. Triple products ( $E_{H_1}E_{H_2}E_{H_1+H_2}$ ) were established such that two of the participating structure factors had  $E$  values of at least 1.30, the third  $E$  value had to be higher than 1.0. Values of the structure invariants were calculated, applying the  $B_{3,0}$  formula (Karle & Hauptman, 1958). The expression used was:

$$E_{H_1} \cdot E_{H_2} \cdot E_{H_1+H_2} = AB + C.$$

$A$  is a positive constant,  $C$  is a small correction term and  $B$  is given by:

$$B = \langle (E_K^2 - 1) (E_{K+H_1}^2 - 1) (E_{K+H_1+H_2}^2 - 1) \rangle_K,$$

in which the mean value is taken over all 1493 reflexions  $K(hkl)$ . Triple products with the highest calculated positive values of  $AB + C$  used were in the evaluation of the structure factor signs by a symbolic procedure executed by hand. One single solution was obtained, and the corresponding Fourier synthesis with 384 terms could easily be interpreted.

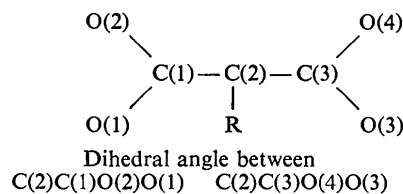
Full-matrix least-squares refinement alternated with

Table 2. Fractional coordinates and thermal parameters of difluoromalonic acid

The  $\beta_{ij}$  values are defined by the expression for the temperature factor:  $\exp(-10^{-4} \sum \sum h_i h_j \beta_{ij})$

	$x$	$y$	$z$	$\beta_{11}$	$\beta_{22}$	$\beta_{33}$	$2\beta_{12}$	$2\beta_{23}$	$2\beta_{31}$
F(1)	0.6471	0.3560	0.1716	101	350	76	17	-56	82
F(2)	0.8027	0.6289	0.3075	112	182	151	-13	60	169
O(1)	0.6819	0.6118	0.5048	92	234	101	-57	-99	113
O(2)	0.5030	0.3599	0.3505	82	223	104	-55	-66	108
O(3)	0.9590	0.2248	0.3507	165	552	242	399	424	307
O(4)	0.8402	0.1175	0.4932	169	298	184	227	238	253
C(1)	0.6242	0.4664	0.3929	66	147	71	21	8	70
C(2)	0.7307	0.4263	0.3140	73	158	81	-6	12	87
C(3)	0.8536	0.2360	0.3966	77	186	92	29	9	96
H(1)	0.6196	0.6176	0.5459	$B = 3 \text{ \AA}^2$					
H(2)	1.0164	0.1633	0.4068	$B = 3 \text{ \AA}^2$					

Table 1. Conformation of some substituted malonic acids



Compound	$\alpha$ -Substituent R	C(1)C(2)R	C(3)C(2)R	Distance O(2)–O(4)	Reference
Methylmalonic acid	–CH <sub>3</sub>	0°	71°	3.03 Å	Derissen (1970)
Aminomalonic acid	–NH <sub>2</sub>	9	18	3.06	Kanters <i>et al.</i> (1966)
Tartronic acid	–OH	15	18	2.92	van Eijck <i>et al.</i> (1965)
Fluoromalonic acid	–F	6	equivalent	2.92	Roelofsen <i>et al.</i> (1971)





The shape and orientation of the vibrational ellipsoids in the difluoromalonic acid molecule are shown in Fig. 2.

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## Structure Cristalline et Moléculaire du Complexe Oestradiol-Propanol

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The crystallographic structure of the complex oestradiol-propanol ( $C_{18}H_{24}O_2 \cdot C_3H_8O$ ) has been determined by X-ray diffraction. Space group:  $P2_12_12_1$ ,  $Z=4$ ,  $a=12.215$  (4),  $b=24.251$  (7),  $c=6.671$  (3) Å. A direct minimization of the residual coefficient  $R$  is used to solve the structure. The conformation of the oestradiol molecule previously described is very slightly perturbed by the presence of alcohol in the crystals.

La conformation de la molécule d'oestradiol:  $C_{18}H_{24}O_2$  a déjà été décrite dans la publication de la structure hémihydratée (Busetta & Hospital, 1972). Pour voir

si l'environnement pouvait affecter la conformation du squelette stéroïdique, nous avons étudié les solvates de l'oestradiol avec le propanol ( $C_3H_8O$ ).

Tableau 1. Coordonnées relatives ( $\times 10^5$ ) et facteurs d'agitation thermique

	<i>x</i>	<i>y</i>	<i>z</i>	$B_{11}$	$B_{22}$	$B_{33}$	$B_{23}$	$B_{13}$	$B_{12}$
C(1)	56106 (31)	12841 (18)	-27030 (60)	0.00689	0.00266	0.01983	-0.00254	-0.00135	-0.00059
C(2)	67189 (32)	12069 (17)	-30300 (64)	0.00745	0.00236	0.02211	-0.00248	0.00078	0.00040
C(3)	74635 (29)	13949 (16)	-16217 (66)	0.00609	0.00203	0.02634	-0.00016	-0.00037	0.00023
C(4)	70847 (32)	16707 (16)	504 (51)	0.00693	0.00215	0.02340	-0.00036	-0.00222	-0.00101
C(5)	59786 (30)	17572 (15)	3847 (61)	0.00709	0.00176	0.02171	-0.00042	-0.00036	-0.00042
C(6)	56369 (33)	20667 (18)	22398 (66)	0.00807	0.00245	0.02325	-0.00344	0.00077	-0.00132
C(7)	44264 (34)	22324 (17)	22608 (73)	0.00842	0.00226	0.02877	-0.00391	0.00629	-0.00061
C(8)	37189 (30)	17511 (15)	15640 (58)	0.00671	0.00200	0.01821	0.00015	0.00082	-0.00007
C(9)	39751 (30)	16233 (16)	-6534 (58)	0.00716	0.00223	0.01790	-0.00015	0.00017	0.00049
C(10)	52094 (30)	15467 (15)	-9965 (60)	0.00671	0.00194	0.01977	0.00053	0.00103	-0.00003
C(11)	32794 (33)	11400 (19)	-14173 (69)	0.00682	0.00309	0.02507	-0.00472	-0.00233	-0.00050
C(12)	20384 (34)	12187 (21)	-10476 (70)	0.00642	0.00379	0.02338	-0.00342	-0.00103	0.00009
C(13)	18038 (30)	13600 (17)	11524 (63)	0.00609	0.00247	0.02108	0.00120	-0.00102	0.00083
C(14)	25069 (31)	18598 (16)	17365 (62)	0.00718	0.00212	0.02071	0.00053	0.00246	0.00085
C(15)	20200 (38)	20504 (20)	37562 (77)	0.00938	0.00288	0.02796	-0.00292	0.00824	-0.00021
C(16)	7819 (37)	19285 (22)	35426 (82)	0.00892	0.00333	0.03141	-0.00180	0.00946	0.00054
C(17)	6616 (32)	16041 (19)	15367 (67)	0.00654	0.00308	0.02465	0.00503	0.00183	0.00106
C(18)	19720 (34)	8584 (19)	24985 (79)	0.00694	0.00252	0.03650	0.00295	-0.00285	-0.00039
O(83)	85730 (21)	13238 (12)	-18576 (48)	0.00654	0.00271	0.03087	-0.00161	0.00021	0.00012
O(97)	-2168 (23)	12213 (15)	16177 (51)	0.00686	0.00422	0.02932	0.00463	0.00112	-0.00143
C(51)	-2915 (95)	-913 (31)	82216 (200)	0.02822	0.00317	0.09413	0.01036	-0.01434	-0.00028
C(52)	-10740 (81)	-1525 (25)	66294 (157)	0.03250	0.00282	0.06586	-0.00104	0.01526	0.00436
C(53)	-8402 (69)	1399 (26)	48036 (111)	0.02607	0.00283	0.04682	-0.00361	0.00842	-0.00318
O(54)	-8668 (30)	7067 (14)	50488 (52)	0.01504	0.00280	0.03293	-0.00041	0.01174	0.00124